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An advanced-level textbook of inorganic chemistry for the graduate (B.Sc) and postgraduate (M.Sc) students of Indian and foreign universities. This book is a part of four volume series, entitled "A Textbook of Inorganic Chemistry - Volume I, II, III, IV". CONTENTS: Chapter 1. Stereochemistry and Bonding in Main Group Compounds: VSEPR theory, $d\pi-p\pi$ bonds, Bent rule and energetic of hybridization. Chapter 2. Metal-Ligand Equilibria in Solution: Stepwise and overall formation constants and their interactions, Trends in stepwise constants, Factors affecting stability of metal complexes with reference to the nature of metal ion and ligand, Chelate effect and its thermodynamic origin, Determination of binary formation constants by pH-metry and spectrophotometry. Chapter 3. Reaction Mechanism of Transition Metal Complexes - I: Inert and labile complexes, Mechanisms for ligand replacement reactions, Formation of complexes from aquo ions, Ligand displacement reactions in octahedral complexes- acid hydrolysis, Base hydrolysis, Racemization of tris chelate complexes, Electrophilic attack on ligands. Chapter 4. Reaction Mechanism of Transition Metal Complexes - II: Mechanism of ligand displacement reactions in square planar complexes, The trans effect, Theories of trans effect, Mechanism of electron transfer reactions - types; Outer sphere electron transfer mechanism and inner sphere electron transfer mechanism, Electron exchange. Chapter 5. Isopoly and Heteropoly Acids and Salts: Isopoly and Heteropoly acids and salts of Mo and W: structures of isopoly and heteropoly anions. Chapter 6. Crystal Structures: Structures of some binary and ternary compounds such as fluorite, antiferite, rutile, antirutile, cristobalite, layer

lattices- CdI₂, BiI₃; ReO₃, Mn₂O₃, corundum, perovskite, Ilmenite and Calcite. Chapter 7. Metal-Ligand Bonding: Limitation of crystal field theory, Molecular orbital theory, octahedral, tetrahedral or square planar complexes, π -bonding and molecular orbital theory. Chapter 8. Electronic Spectra of Transition Metal Complexes: Spectroscopic ground states, Correlation and spin-orbit coupling in free ions for 1st series of transition metals, Orgel and Tanabe-Sugano diagrams for transition metal complexes (d¹ - d⁹ states), Calculation of Dq, B and β parameters, Effect of distortion on the d-orbital energy levels, Structural evidence from electronic spectrum, John-Teller effect, Spectrochemical and nephelauxetic series, Charge transfer spectra, Electronic spectra of molecular addition compounds. Chapter 9. Magnetic Properties of Transition Metal Complexes: Elementary theory of magneto-chemistry, Gouy's method for determination of magnetic susceptibility, Calculation of magnetic moments, Magnetic properties of free ions, Orbital contribution, effect of ligand-field, Application of magneto-chemistry in structure determination, Magnetic exchange coupling and spin state cross over. Chapter 10. Metal Clusters: Structure and bonding in higher boranes, Wade's rules, Carboranes, Metal Carbonyl Clusters - Low Nuclearity Carbonyl Clusters, Total Electron Count (TEC). Chapter 11. Metal- π Complexes: Metal carbonyls, structure and bonding, Vibrational spectra of metal carbonyls for bonding and structure elucidation, Important reactions of metal carbonyls; Preparation, bonding, structure and important reactions of transition metal nitrosyl, dinitrogen and dioxygen complexes; Tertiary phosphine as ligand.

Metal clusters are on the brink between molecules and nanoparticles in size. With molecular, nano-scale, metallic as well as non-metallic aspects, metal clusters are a growing,

interdisciplinary field with numerous potential applications in chemistry, catalysis, materials and nanotechnology. This third volume in the series of hot topics from inorganic chemistry covers all recent developments in the field of metal clusters, with some 20 contributions providing an in-depth view. The result is a unique perspective, illustrating all facets of this interdisciplinary area: * Inter-electron Repulsion and Irregularities in the Chemistry of Transition Series * Stereochemical Activity of Lone Pairs in Heavier Main Group Element Compounds * How Close to Close Packing? * Forty-Five Years of Praseodymium Diiodide * Centered Zirconium Clusters * Titanium Niobium Oxichlorides * Trinuclear Molybdenum and Tungsten Cluster Chalcogenides * Current State of (B,C,N)-Compounds of Calcium and Lanthanum * Ternary Phases of Lithium with Main-Group and Late-Transition Metals * Polar Intermetallics and Zintl Phases along the Zintl Border * Rare Earth Zintl Phases * Structure-Property Relationships in Intermetallics * Ternary and Quaternary Niobium Arsenide Zintl Phases * The Building Block Approach to Understanding Main-Group-Metal Complex Structures * Cation-Deficient Quaternary Thiospinels * A New Class of Hybrid Materials via Salt Inclusion Synthesis * Layered Perrhenate and Vanadate Hybrid Solids * Hydrogen Bonding in Metal Halides * Syntheses and Catalytic Properties of Titanium Nitride Nanoparticles * Solventless Thermolysis * New Potential Scintillation Materials in Borophosphate Systems. With its didactical emphasis, this volume addresses a wide readership, such that both students and specialists will profit from the expert contributions.

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"This book has succeeded in covering the basic chemistry essentials required by the pharmaceutical science student—the undergraduate reader, be they chemist, biologist or pharmacist will find this an interesting and valuable read."—Journal of Chemical Biology, May 2009

Chemistry for Pharmacy Students is a student-friendly introduction to the key areas of chemistry required by all pharmacy and pharmaceutical science students. The book provides a comprehensive overview of the various areas of general, organic and natural products chemistry (in relation to drug molecules). Clearly structured to enhance student understanding, the book is divided into six clear sections. The book opens with an overview of general aspects of chemistry and their importance to modern life, with particular emphasis on medicinal applications. The text then moves on to a discussion of the concepts of atomic structure and bonding and the fundamentals of stereochemistry and their significance to pharmacy— in relation to drug action and toxicity. Various aspects of aliphatic, aromatic and heterocyclic chemistry and their pharmaceutical importance are then covered with final chapters looking at organic reactions and their applications to drug discovery and development and natural products chemistry.

accessible introduction to the key areas of chemistry required for all pharmacy degree courses student-friendly and written at a level suitable for non-chemistry students includes learning objectives at the beginning of each chapter focuses on the physical properties and actions of drug molecules

The first IUPAC Manual of Symbols and Terminology for Physicochemical Quantities and Units (the Green Book) of which this is the direct successor, was published in 1969, with the object of 'securing clarity and precision, and wider agreement in the use of symbols, by chemists in

different countries, among physicists, chemists and engineers, and by editors of scientific journals'. Subsequent revisions have taken account of many developments in the field, culminating in the major extension and revision represented by the 1988 edition under the simplified title *Quantities, Units and Symbols in Physical Chemistry*. This 2007, Third Edition, is a further revision of the material which reflects the experience of the contributors with the previous editions. The book has been systematically brought up to date and new sections have been added. It strives to improve the exchange of scientific information among the readers in different disciplines and across different nations. In a rapidly expanding volume of scientific literature where each discipline has a tendency to retreat into its own jargon this book attempts to provide a readable compilation of widely used terms and symbols from many sources together with brief understandable definitions. This is the definitive guide for scientists and organizations working across a multitude of disciplines requiring internationally approved nomenclature.

This part of the book deals mainly with chemistry of the elements meant for 3rd year degree students.

Dr. Alan Williams has acquired a considerable experience in work with transition metal complexes at the Universities of Cambridge and Geneva. In this book he has tried to avoid the variety of ephemeral and often contradictory rationalisations encountered in this field, and has made a careful comparison of modern opinions about chemical bonding. In my opinion this effort is fruitful for all students and active scientists in the field of inorganic chemistry. The

distant relations to group theory, atomic spectroscopy and epistemology are brought into daylight when Dr. Williams critically and pedagogic ally compares quantum chemical models such as molecular orbital theory, the more specific L. C. A. O. description and related "ligand field" theory, the valence bond treatment (which has conserved great utility in antiferromagnetic systems with long inter nuclear distances), and discusses interesting, but not too well-defined concepts such as electronegativity (also derived from electron transfer spectra), hybridisation, and oxidation numbers. The interdisciplinary approach of the book shows up in the careful consideration given to many experimental techniques such as vibrational (infra-red and Raman), electronic (visible and ultraviolet), Mossbauer, magnetic resonance, and photoelectron spectra, with data for gaseous and solid samples as well as selected facts about solution chemistry. The book could not have been written a few years ago, and is likely to remain a highly informative survey of modern inorganic chemistry and chemical physics. Geneva, January 1979 C. K.

While existing books related to DOE are focused either on process or mixture factors or analyze specific tools from DOE science, this text is structured both horizontally and vertically, covering the three most common objectives of any experimental research: * screening designs * mathematical modeling, and * optimization. Written in a simple and lively manner and backed by current chemical product studies from all around the world, the book elucidates basic concepts of statistical methods, experiment design and optimization techniques as applied to chemistry and chemical engineering. Throughout, the focus is on unifying the theory and methodology of optimization with well-known statistical and experimental methods. The author

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draws on his own experience in research and development, resulting in a work that will assist students, scientists and engineers in using the concepts covered here in seeking optimum conditions for a chemical system or process. With 441 tables, 250 diagrams, as well as 200 examples drawn from current chemical product studies, this is an invaluable and convenient source of information for all those involved in process optimization.

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