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The final mass of copper

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(49.5g Cu) ended up significantly more than the original value (1.962 Cu). The final moles of copper (.77 moles Cu) ended up being significantly more than the initial moles of copper (.03 moles Cu). And

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the percent yield of copper ended up being 2556.67 percent which is extremely high.

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Throughout this lab, the

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same sample of copper formed various different compounds, from copper hydroxide to copper (II) oxide. During all of these reactions, the mass of copper remained constant, for the Law of Conservation of Mass states

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it so. Stoichiometry can be used to illustrate how the mass remains constant during the experiment.

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Lab 31 Answer Key~~

In the lab, the copper was dissolved in nitrate acid which released a brown smoke

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and the liquid turned a pure blue. Then, the beaker was put in an ice bath and added sodium hydroxide in order to change the state to a solid. It was then headed to separate the solid from the liquid. It was decanted to

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Lab - Yamilet's AP Chemistry
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Stoichiometry Using Copper
Purpose: The purpose is to
see how the amount of copper

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(and copper itself) is
altered after a series of
reactions.

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The initial mass of copper
was 2.003 grams. The final
mass of copper was 9.256

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grams of copper. The initial and final masses of copper are supposed to be the same, but they are different. The initial moles of copper is 0.03152 mol, and the final moles of copper is 0.1457 mol.

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~~Copper Lab — AP Chemistry~~

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LAB 1 Stoichiometry Using
Copper Lab Lauren Rogers
Second Period AP Chemistry
STOICHIOMETRY USING COPPER

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LAB 2 Purpose: The purpose of the experiment was to observe how copper was affected by a series of chemical reactions to prove that copper was able to be recovered and maintain its integrity.

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Stoichiometry Using Copper
Lab. October 23, 2012.

Purpose. A solid copper
metal of known mass is
performed with a series of

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reactions, eventually recovering the copper at the end and testing the Law of Conservation of Mass. Quantitative Data.

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some students is that the
individual particles of a

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a formula of that material.
STOICHIOMETRY: The Reaction
of Iron with Copper (II)
Sulfate

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Stoichiometry Lab. In this experiment, you will decompose a mixture of basic copper II carbonate [with the formula $\text{CuCO}_3 \cdot \text{Cu}(\text{OH})_2$] to form copper II oxide, carbon dioxide and water. You will determine

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the moles of reactant used
and product produced through
careful measurement of
masses and by stoichiometry.

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In this experiment, iron is

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more active than copper.

Iron forms 2 types of ions, namely Fe^{2+} and Fe^{3+} . We shall use stoichiometric principles to determine which of these ions is formed in the reaction between iron and copper(II)

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~~General Chemistry I (FC, 09
— 10) Lab #4: Stoichiometry~~

~~...~~

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Calculate the percent yield of the copper in REACTION 1 and of the carbon dioxide in REACTION 2 using the equation below (show your work). % yield =
experimental yield (100

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theoretical yield. A perfect percent yield would be 100%. For each reaction, comment on your degree of accuracy and suggest possible sources of measurement error.

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Copper/Iron Stoichiometry

Grace Timler AB1 October 3,
2017 Abstract The techniques
used in this lab are
quantitative transfer and
vacuum filtration with the
reaction of 8.001 grams of
copper (II) sulfate, CuSO_4 ,

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and 2.0153 grams of iron powder, Fe. The goal of this experiment was to determine the product of copper (II) sulfate with iron.

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This work details minor, trace and ultratrace methods; addresses the essential stages that precede measurement; and highlights the measurement

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Systems most likely to be used by the pragmatic analyst. It features key material on inclusion and phase isolation. The book is designed to provide useful maps and signposts for metals analysts who must

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Verify that stringent trace level compositional specifications have been met.

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Steve and Susan Zumdahl's
texts focus on helping
students build critical
thinking skills through the

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process of becoming independent problem-solvers. They help students learn to think like a chemists so they can apply the problem solving process to all aspects of their lives. In CHEMISTRY: AN ATOMS FIRST

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APPROACH, the Zumdahls use a meaningful approach that begins with the atom and proceeds through the concept of molecules, structure, and bonding, to more complex materials and their properties. Because this

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approach differs from what most students have experienced in high school courses, it encourages them to focus on conceptual learning early in the course, rather than relying on memorization and a plug

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and chug method of problem solving that even the best students can fall back on when confronted with familiar material. The atoms first organization provides an opportunity for students to use the tools of critical

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to apply rules and models
and to evaluate outcomes.
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